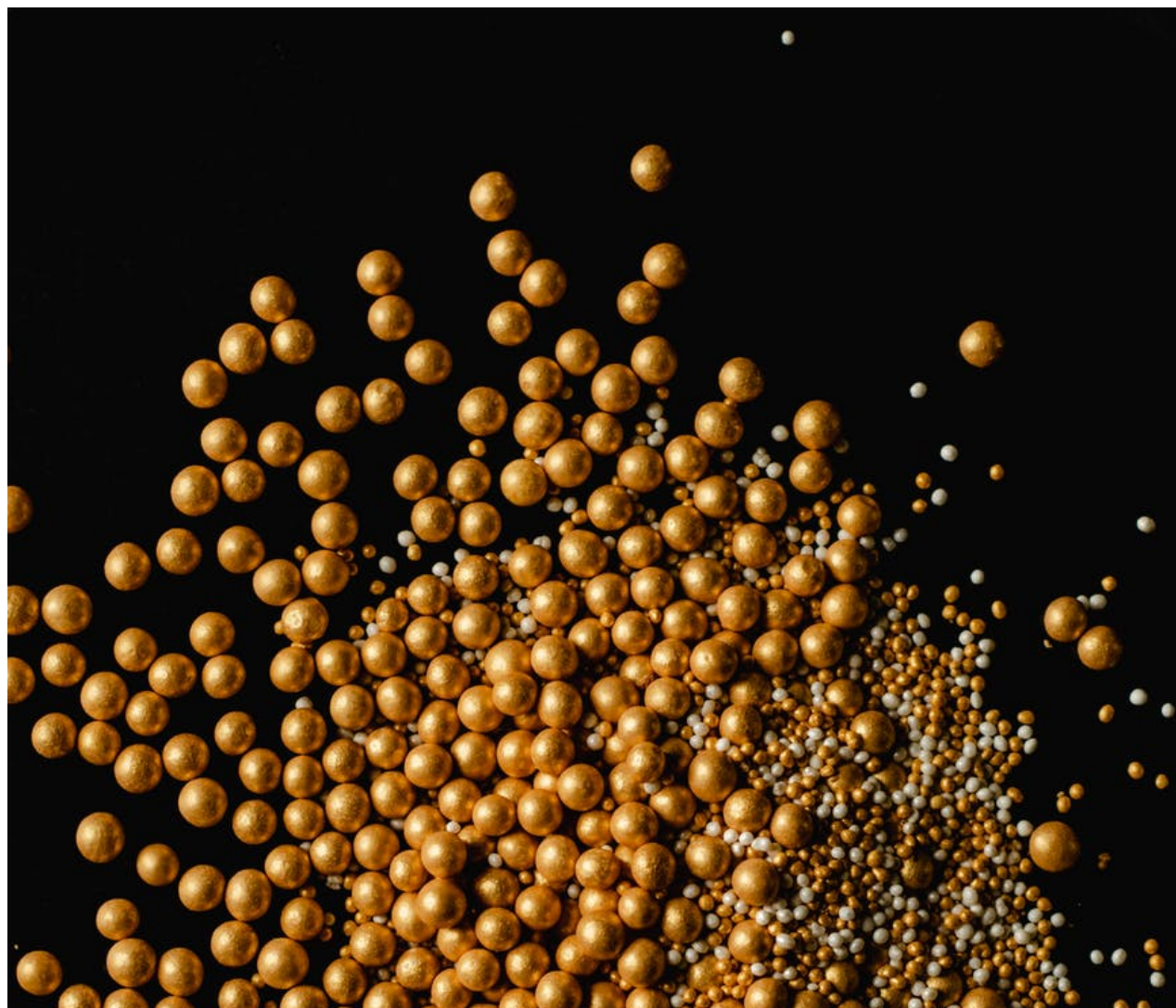
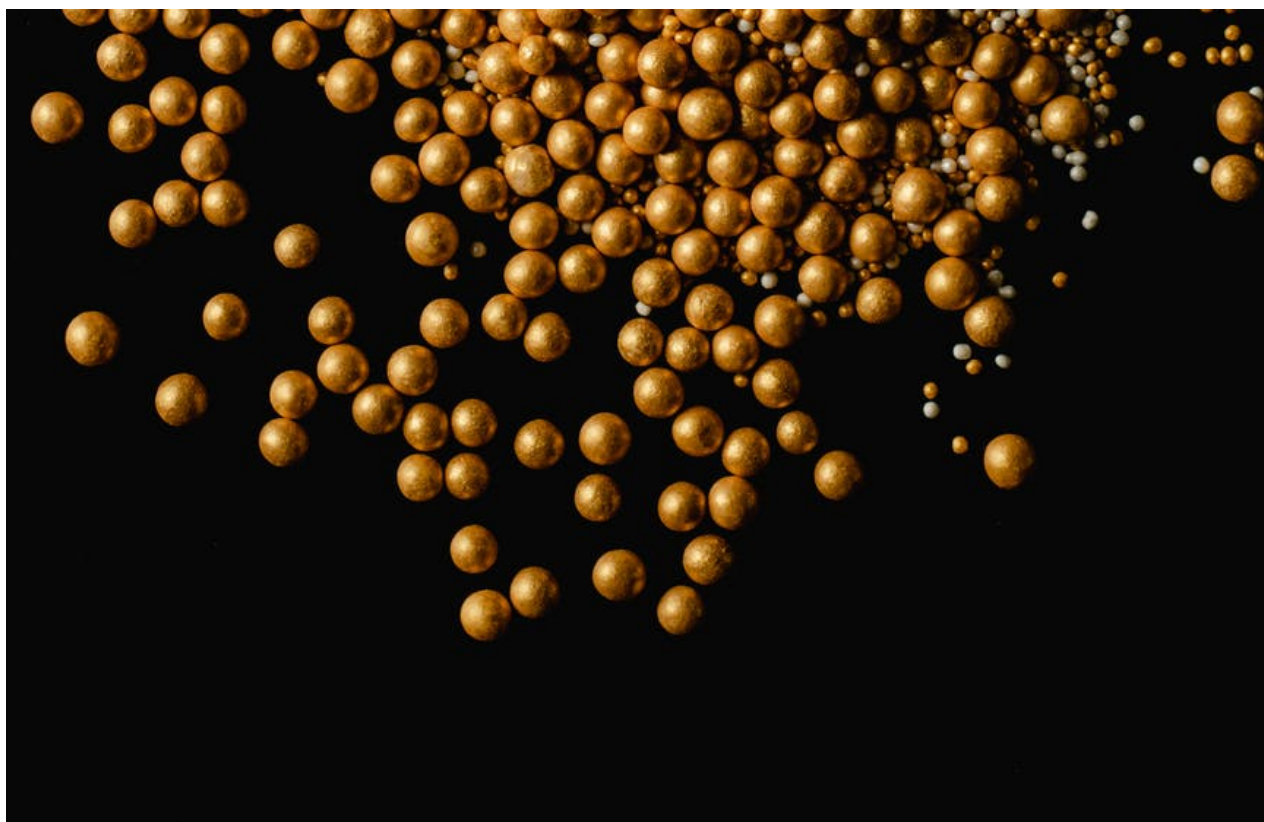




Priya's Learning Centre

Ch – 4 Structure Of The Atom (Part 3)





Q.1: _____ are defined as the atoms of the same element, having the same atomic number, but different mass numbers.

- a) Isobars
- b) Isotopes
- c) Both

Q.2: Each _____ of an element is a pure substance.

- a) Isobar
- b) Isotope
- c) Both

Q.3: An isotope of _____ is used in the treatment of cancer.

- a) hydrogen

b) cobalt

c) iodine

Q.4: The mass of an _____ of any natural element is taken as the average mass of all the naturally occurring atoms of that element.

a) electron

b) atom

c) none

Q.5: Atoms of different elements with different atomic numbers but same mass number are known as _____.

a) Isotopes

b) Isobars

c) None

Q.6: Which of the following statements are true ?

i) The shells of an atom are categorised as K,L,M,N...

ii) The three isotopes of a hydrogen atom are protium, deuterium and tritium.

iii) Isobars and Isotopes are same.

a) i and iii

b) i and ii

c) ii and iii

Q.7: An isotope of _____ is used to treat goitre.

a) iron

b) iodine

c) zinc

Q.8: Elements are defined by the number of _____ they possess.

a) electrons

b) protons

c) neutrons

Q.9: Which of the following element's valency is 0 ?

a) Lithium

b) Neon

c) Aluminium

Q.10: Which of the following element's atomic number is 15 ?

a) Argon

b) Phosphorous

c) Barium

Answers:

1. b
2. b
3. b
4. b
5. b
6. b
7. b
8. b
9. b
10. b

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